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# Secondary Organic Aerosol Formation from Aromatic Alkene Ozonolysis: Influence of the Precursor Structure on Yield, Chemical Composition, and Mechanism

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3  
4 **Influence of the Precursor Structure on Yield, Chemical Composition and**  
5  
6 **Mechanism**  
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## Abstract

The influence of the precursor chemical structure on secondary organic aerosol (SOA) formation was investigated through the study of the ozonolysis of two anthropogenic aromatic alkenes: 2-methylstyrene and indene. Experiments were carried out in three different simulation chambers: ICARE 7300L FEP Teflon chamber (ICARE, Orléans, France), EUPHORE FEP Teflon chamber (CEAM, Valencia, Spain) and CESAM evacuable stainless steel chamber (LISA, Créteil, France). For both precursors, SOA yield and growth were studied on a large range of initial concentrations (from ~60 ppbv to 1.9 ppmv) and the chemical composition of both gaseous and particulate phases was investigated at a molecular level. Gas-phase was described using FTIR spectroscopy and on-line gas chromatography coupled to mass spectrometry and particulate chemical composition was analysed i) on-line by thermo-desorption coupled to chemical ionisation mass spectrometry (TD-API-AMS) and ii) off-line by supercritical fluid extraction coupled to gas chromatography and mass spectrometry (SFE-GC-MS). The results obtained from a large set of experiments performed in three different chambers and using several complementary analytical techniques were in very good agreement. SOA yield was up to 10 times higher for indene ozonolysis than for 2-methylstyrene ozonolysis at the same reaction advancement. For 2-methylstyrene ozonolysis, formaldehyde and o-tolualdehyde were the two main gaseous phase products while o-toluic acid was the most abundant amongst six products detected within the particulate phase. For indene ozonolysis, traces of formic and phthalic acids, as well as eleven species were detected in the gaseous phase and eleven other products were quantified in the particulate phase, where phthaldialdehyde was the main product. Based on the identified products, reaction mechanisms were proposed that highlight specific pathways due to the precursor chemical structure. These mechanisms were finally compared and discussed regarding SOA formation. In the case of 2-methylstyrene ozonolysis, ozone adds mainly on the external and mono-substituted double bond, yielding only one C<sub>8</sub>- and mono-functionalized Criegee intermediate and hence more volatile products as well as lower SOA mass than indene ozonolysis in similar experimental conditions. In the case of indene, ozone adds mainly on the five-carbon-ring and di-substituted C=C double bond, leading to the

1  
2 formation of two C9- and bi-functionalized Criegee intermediates, which then evolve *via*  
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4 different pathways including the hydroperoxide channel and form highly condensable first-  
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6 generation products.  
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## 12 **1. Introduction**

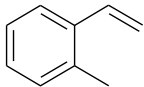
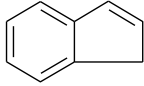
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14 Airborne particles influence visibility, cloud formation, light absorption and scattering, and  
15 thus they play a key role on climate (IPCC, 2013).<sup>1</sup> Particle impacts on health have been  
16 observed and confirmed for many years now by epidemiologic studies showing the influence  
17 of their size.<sup>2-5</sup> Secondary organic aerosol (SOA) is formed in the atmosphere by gas-particle  
18 partitioning of semi- or low volatile oxidation products of volatile organic compounds (VOCs)  
19 and accounts for a significant fraction of tropospheric fine aerosol, in remote areas as well as  
20 in urban areas.<sup>6,7</sup>  
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29 It is critical for atmospheric modelling to describe SOA formation and understand processes  
30 at the molecular level since aerosol impacts on health and climate are tightly linked to  
31 chemical composition. The need for reliable mechanisms describing SOA formation was  
32 reported by landmark reviews<sup>8-11</sup> and more recent papers, as only incomplete or  
33 oversimplified schemes are yet available.<sup>12-15</sup> This implies the investigation of the parameters  
34 influencing SOA formation. The precursor molecular structure, since it controls the oxidation  
35 mechanisms and hence the nature of the secondary products, is one of these key  
36 parameters. As an example, Bonn et al. observed that aerosol formation yields highly  
37 depend on the location of the C=C double bond: SOA yields are higher for  $\alpha$ -pinene  
38 (endocyclic double bond) than for  $\beta$ -pinene (exocyclic double bond) ozonolysis.<sup>16</sup> A few other  
39 studies have also investigated the influence of the hydrocarbon structure on SOA formation  
40 yields. Keywood et al. investigated a series of cycloalkenes ozonolysis and found that (i)  
41 SOA yields increase with increasing number of carbons in the cycloalkane ring, (ii) SOA  
42 yields are enhanced by the presence of a methyl group located at a C=C double-bonded site  
43 but reduced by the presence of a methyl group at a non-double-bonded site.<sup>17</sup> Ng et al.  
44 reported that, for a large set of terpenes, regardless of the reaction systems (ozonolysis or  
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photooxidation), precursors can be classified into two groups : (i) compounds with only one C=C double bond, for which the first oxidation step is rate-limiting and aerosols are formed mainly from low volatility first-generation oxidation products and (ii) compounds with more than one double bond, for which the second oxidation step may also be rate-limiting and second-generation products contribute substantially to SOA growth.<sup>18</sup>

To further investigate the influence of the precursor structure on SOA formation arising from anthropogenic precursors, this study focuses on the ozonolysis of two mono-aromatic compounds bearing an additional external and conjugated unsaturation available to react with ozone: 2-methylstyrene and indene (Table 1). These two compounds both exhibit very similar structures, they only differ by the structural configuration of the reactive C=C double bond which is involved in a cycle (indene) or not (2-methylstyrene). These two VOCs were detected in the atmosphere<sup>19</sup> and were found to be emitted by biomass burning<sup>20-23</sup> and automobile exhausts.<sup>24,25</sup>

**Table 1. Molecular Structures and Reactivity Rate Constants of the Aromatic Alkenes**

Parent hydrocarbon	Molecular structure	Formula MW g mol <sup>-1</sup>	k <sub>O<sub>3</sub></sub> cm <sup>3</sup> molecule <sup>-1</sup> s <sup>-1</sup>	k <sub>OH</sub> cm <sup>3</sup> molecule <sup>-1</sup> s <sup>-1</sup>
2-methylstyrene		C <sub>9</sub> H <sub>10</sub> 118	(3.1 ± 0.6) 10 <sup>-17</sup> (26)	-
indene		C <sub>9</sub> H <sub>8</sub> 116	(1.6 ± 0.4) 10 <sup>-16</sup> (26) (1.7 ± 0.5) 10 <sup>-16</sup> (27)	7.8 10 <sup>-11</sup> (27)

This work focuses for the first time on the formation of SOA from the ozonolysis of 2-methylstyrene and indene and investigating the role of their structure and product formation pathways on SOA formation. The experiments were performed using three different simulation chambers: ICARE 7300L FEP Teflon chamber (ICARE, Orléans, France), EUPHORE FEP Teflon chamber (CEAM, Valencia, Spain) and CESAM evacuable stainless steel chamber (LISA, Créteil, France). The interest of the present work lies in the description of the gaseous and particulate phase chemical composition at the molecular level and in the

1  
2 reaction pathways proposed for each compound identified in the aerosol. The comparison of  
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4 ozonolysis mechanisms for 2-methylstyrene and indene allows for discussion of aerosol  
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6 formation and chemical composition related to the precursor chemical structure.  
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## 10 **2. Experimental section**

### 11 *2.1. Chemicals and gases*

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14 2-Methylstyrene (95 % purity), indene (99 % purity), o-tolualdehyde (99 % purity), benzoic  
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16 acid (> 99.5 % purity), phenyl formate (95 % purity), 2-acetylbenzoic acid (98 % purity), 3-  
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18 hydroxybenzaldehyde (> 97 % purity), phthalic acid (99 % purity), 3-hydroxy benzoic acid (99  
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20 % purity) and BSTFA (N,O bis(trimethylsilyl)trifluoroacetamide containing 1%  
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22 trimethylchlorosilane) were purchased from Sigma Aldrich. Tridecane (purity > 97 %) and 2-  
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24 methylphenol (> 99 %) and 1,3-dioxolane (99%) were purchased from Fluka. Sabinaketone  
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26 was synthesized in MEREIA group in LISA (Créteil, France).  
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29 Purified air was used for ICARE and EUPHORE chamber experiments, whereas a mixture of  
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31 N<sub>2</sub> produced from evaporation of pressurized liquid nitrogen (Air Liquide) and O<sub>2</sub> (Alphagaz  
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33 1, Air Liquide) was used for CESAM chamber experiments. Helium (alphagaz 1), methane  
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35 (N45), SF<sub>6</sub> (0.9997), and CO<sub>2</sub> (SFE-grade) were purchased from Air Liquide.  
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37 Dichloromethane (HPLC grade) was purchased from Prolabo.  
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### 42 *2.2. Simulation chamber experiments and SOA production*

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44 In order to account for possible wall effects on SOA yields or composition and to take  
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46 advantage of the diversity of various analytical environments, three different chambers were  
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48 used for this study. Two of them, ICARE and EUPHORE chambers, are made of FEP Teflon  
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50 film (with surface to volume (S/V) ratios of 3.2 and 0.66 m<sup>-1</sup> respectively) and the third one is  
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52 a stainless steel evacuable chamber, CESAM (S/V = 4.3 m<sup>-1</sup>). The main characteristics of  
53  
54 each chamber are provided in Table 2. A more complete description of these installations  
55  
56 has been previously reported in the literature by Le Person et al. for the ICARE chamber<sup>26</sup>,  
57  
58 Klotz et al. and Martin-Reviejo and Wirtz, for the EUPHORE chamber<sup>28,29</sup> and Wang et al. for  
59  
60 the CESAM chamber.<sup>30</sup>

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2 In the following paragraphs, 2-methylstyrene and indene are referred to as MeS and Ind.  
3  
4 ICARE, EUPHORE and CESAM chambers are referred to as I, E and C, respectively.  
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6 Hence, experiments are named according to the precursor, the chamber and the number of  
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8 the experiment (e.g. Ind\_E1 is experiment # 1 performed in EUPHORE chamber with indene  
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10 as precursor).

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12 All instruments used for gaseous phase and particulate phase characterisation are listed in  
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14 Table 2. Gases were analysed using ozone UV absorption monitor, FTIR spectroscopy as  
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16 well as on-line gas chromatography coupled to mass spectrometry (GC-MS) and off-line  
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18 liquid chromatography coupled to mass spectrometry (LC-MS) after sampling with DNPH  
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20 cartridges for carbonyl compound analysis specifically (see SI1 for details). Gas phase  
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22 product yields were obtained by plotting product concentrations as a function of consumed  
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24 parent hydrocarbon (HC) ( $[products] = f(\Delta[HC])$ ). Hence, average product yields could be  
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26 determined from the linear least-square regression of experimental data. The particulate  
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28 phase was physically characterized using Scanning Mobility Particle Sizer (SMPS) and a  
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30 Tapered Element Oscillating Microbalance monitor (TEOM) and chemically described by on-  
31  
32 line thermo-desorption – Atmospheric Pressure Ionisation – Aerosol Mass spectrometry (TD-  
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34 API-AMS)<sup>31</sup> and off-line supercritical Fluid Extraction coupled to Gas Chromatography and  
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36 Mass spectrometry (SFE-GC-MS) after sampling on filters (GF/F, 47 mm, Whatman).<sup>32</sup>

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38 The amount of reacted hydrocarbon varied from ~ 60 ppbv to ~ 1.9 ppmv for 2-methylstyrene  
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40 and from ~ 80 ppbv and ~ 0.9 ppmv for indene. All experiments were performed under dry  
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42 conditions (%RH < 5%). No radical scavenger or seed aerosol was used. During MeS\_I3 and  
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44 Ind\_I2 experiments only, 1,3-dioxolane (2 ppmv) was used as OH tracer. Experimental  
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46 conditions are reported in Table 2, as well as the order of reactant introduction. More details  
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48 on SOA formation experimental procedures and instrument parameters are provided in the  
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50 Supporting Information (SI2.).  
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**Table 2. Initial Experimental Conditions and SOA Formation Yields**

Simulation chamber	Parent hydrocarbon	Exp #	[HC] <sub>0</sub> (ppbv)	[O <sub>3</sub> ] <sub>0</sub> (ppbv)	T (K)	Δ[HC] (μg m <sup>-3</sup> )	ΔM <sub>0</sub> (μg m <sup>-3</sup> )	Y <sub>SOA</sub> (%)
ICARE (a), (d), (e)	2-methylstyrene	MeS_I1	900 <sup>(g)</sup>	870	298±2	4340±250	-	-
ICARE (a), (d), (e)	2-methylstyrene	MeS_I2	1900 <sup>(g)</sup>	820	298±2	5020±250	-	-
ICARE (a), (d), (e)	2-methylstyrene	MeS_I3	1700 <sup>(g)</sup>	670 <sup>(h)</sup>	296±2	4530±250	-	-
CESAM (c)	2-methylstyrene	MeS_C1	948 <sup>(f)</sup>	58	294±1	170±20	2.6±0.5	1.5±0.2
CESAM (c)	2-methylstyrene	MeS_C2	948 <sup>(f)</sup>	148 <sup>(i)</sup>	294±1	580±20	14.2±0.5	2.4±0.1
CESAM (c)	2-methylstyrene	MeS_C3	957 <sup>(f)</sup>	105	296±2	417±20	12.1±0.5	2.9±0.2
CESAM (c)	2-methylstyrene	MeS_C4	957 <sup>(f)</sup>	193 <sup>(i)</sup>	296±2	950±20	21.8±0.5	2.3±0.1
CESAM (c)	2-methylstyrene	MeS_C5	957 <sup>(f)</sup>	413 <sup>(i)</sup>	296±2	1941±20	54.5±0.5	2.8±0.1
ICARE (a), (d), (e)	indene	Ind_I1	800 <sup>(g)</sup>	950	297±2	3460±200	-	-
ICARE (a), (d), (e)	indene	Ind_I2	900 <sup>(g)</sup>	950	298±2	3900±200	-	-
EUPHORE (b), (d)	indene	Ind_E1	756 <sup>(f)</sup>	>1000	294±1	3900±200	3750±200	96±5
EUPHORE (b), (d)	indene	Ind_E2	100 <sup>(g)</sup>	124	293±1	464±30	101±6	22±3
EUPHORE (b), (d)	indene	Ind_E3	100 <sup>(g)</sup>	118	304±2	450±30	54±3	12.1±1.5
CESAM (c)	indene	Ind_C1	789 <sup>(f)</sup>	270	296±1	1686±20	462±0.5	27.4±0.3
CESAM (c)	indene	Ind_C2	789 <sup>(f)</sup>	26 <sup>(i)</sup>	297±1	113±20	10±0.5	8.8±0.8
CESAM (c)	indene	Ind_C3	789 <sup>(f)</sup>	100 <sup>(i)</sup>	297±1	454±20	69±0.5	15.2±0.4
CESAM (c)	indene	Ind_C4	789 <sup>(f)</sup>	281 <sup>(i)</sup>	297±1	1248±20	314±0.5	25.1±0.3
CESAM (c)	indene	Ind_C5	778 <sup>(f)</sup>	70	294±1	527±20	90±0.5	17.1±0.5
CESAM (c)	indene	Ind_C6	778 <sup>(f)</sup>	140 <sup>(i)</sup>	294±1	895±20	137±0.5	15.3±0.3

(a) indoor, FEP Teflon chamber, V = 7.3 m<sup>3</sup>, S/V = 3.2 m<sup>-1</sup>, O<sub>3</sub> UV absorption monitor (APOA-360, Horiba), FTIR (Magna 5700, Nicolet), SMPS: CPC (3022A, TSI) + DMA (3081, TSI).<sup>26</sup>

(b) outdoor + retractable steel housing closed during experiments, FEP Teflon, V = 200 m<sup>3</sup>, S/V = 0.66 m<sup>-1</sup>, O<sub>3</sub> UV absorption monitor (ML9810, Monitor Labs), FTIR (Magna 550, Nicolet), on-line GC-PID (8130, Fisons), on-line GC-MS (Saturn 2000, Varian), off-line LC-MS (LCQ Advantage Thermo Finnigan) after sampling on DNPH cartridges, SMPS: CPC (3010, TSI) + DMA (3081, TSI), TEOM (1400a, Rupprecht and Patashnick) operated : at 27°C, 1 min accumulation time and using PM1 sampling head.<sup>28, 29</sup>

(c) indoor, stainless steel, V = 4.2 m<sup>3</sup>, S/V = 4.3 m<sup>-1</sup>, O<sub>3</sub> UV absorption monitor (APOA-370, Horiba), FTIR (Tensor 37, Bruker), SMPS: CPC (3010, TSI) + DMA (3080, TSI).<sup>30</sup>

(d) on-line TD-API-AMS.<sup>31</sup>

(e) off-line SFE-GC-MS after sampling on filters (GF/F, 47mm, Whatman).<sup>32</sup>

(f) The parent hydrocarbon was introduced first. In Ind\_E1 only, ozone was added slowly in order to decrease the reaction rate.

(g) Ozone was introduced before the parent hydrocarbon.

(h) Experiment was performed in presence of 1,3-dioxolane (2 ppmv) used as OH tracer.

(i) Experiment was performed just following the previous one without chamber evacuation. The parent hydrocarbon was introduced in large excess and small quantities of ozone were added in the chamber once gas-particle equilibrium from the previous experiment was established and aerosol growth completed.



### 2.3. Aerosol yield and aerosol growth

For each experiment, SOA yield ( $Y$ ) was calculated from the ratio of the formed aerosol mass concentration ( $\Delta M_0$ ) to the amount of reacted hydrocarbon ( $\Delta HC$ ) at the maximum particle concentration. The aerosol mass concentration was measured directly from the TEOM instrument when available or calculated using SMPS measured volume concentration assuming a specific SOA density value. Density measurements arising from the parallel use of TEOM and SMPS ranged from 0.8 to 1.2 g cm<sup>-3</sup>. Temperature is one varying parameter from one experiment to the other (Table 2) which could contribute to the observed differences in density. Nevertheless, density estimations can only be considered as qualitative here and should be determined more precisely with complementary experiments. Therefore, a density value of 1 g cm<sup>-3</sup> was assumed for SOA mass calculation from aerosol volume provided by SMPS measurements. Parent hydrocarbon and particle concentrations were corrected for dilution and wall losses (i) by applying a first order loss rate obtained from the measured decay of the particles at the end of each experiment, in EUPHORE experiments or (ii) based on the flow rate used for chamber inner pressure regulation in the CESAM chamber, which exhibits negligible wall losses in this particle size range and within the experiment timescale.<sup>30</sup>

For each hydrocarbon, “yield curves” were plotted by fitting experimental yields versus final  $\Delta M_0$  using the volatility basis set (VBS) model developed by Donahue and co-workers, further referred to as the four-product model and described by the following equation :

$$Y = \sum_i \frac{\alpha_i}{1 + C_i^* / \Delta M_0}$$

where  $C_i^*$  is the saturation concentration of compound  $i$ .<sup>34</sup> For parametrisation, experimental data points were fitted using non-linear least square regressions.

SOA time-dependent growth curves, i.e.  $\Delta M_0$  as a function of the amount of hydrocarbon reacted  $\Delta[HC]$ , were plotted for all experiments, in order to investigate the relative contribution of first- and second-generation products to SOA formation and growth.<sup>18</sup>

### 3. Results

#### 3.1. Gas-phase chemical composition

##### 2-Methylstyrene ozonolysis

The gas phase analysis by FTIR spectroscopy showed the formation of formaldehyde, formic acid and o-tolualdehyde. The main products, formaldehyde ( $52 \pm 5$  %) and o-tolualdehyde ( $50 \pm 5$  %) were found to be of primary origin. Their formation yields were determined for the five experiments performed in the ICARE and the CESAM chambers (Table 3). Formaldehyde and o-tolualdehyde formation yields were higher in CESAM chamber and the sum of both products was slightly exceeding 100%. This may be explained by either underestimation of uncertainties or a small secondary contribution for these products possibly due to wall effects. Formic acid was produced in far lower quantities (about 2%) and was probably a secondary reaction product. These results are consistent and comparable with previous results reported for styrene ozonolysis where formaldehyde and benzaldehyde were the major products formed with yields of  $\sim 37 \pm 5$  % and  $41 \pm 5$  % respectively<sup>35</sup> and  $35 \pm 5$  % for both compounds.<sup>36</sup> Formic acid was also observed previously as a minor product with yields of about 1-2 %.<sup>35</sup>

Toluic acid, which was the main component detected in the particulate phase, was sought in the gas phase but not detected.

**Table 3. Molar Yields of the Gas Phase Products Formed during 2-Methylstyrene Ozonolysis**

Experiments (#)	Gaseous product yields (%)		
<i>2-Methylstyrene ozonolysis</i>	<i>Formaldehyde</i>	<i>o-Tolualdehyde</i>	<i>Formic acid</i>
MeS_I1	$55 \pm 4$	$55 \pm 5$	$\sim 2$
MeS_I2	$39 \pm 5$	$41 \pm 5$	$\sim 1$
MeS_I3	$51 \pm 6$	$50 \pm 5$	$\sim 2$
MeS_C2	$59 \pm 5$	$59 \pm 5$	(a)
MeS_C3	$54 \pm 5$	$58 \pm 5$	(a)
Average mean	$52 \pm 5$	$50 \pm 5$	$< 2$

(a) not detected.

### Indene ozonolysis

In ICARE chamber experiments, formic acid was detected in the gas phase at very low concentrations, below the quantification limit (about 25 ppb defined as 10 times the standard deviation). Phthalic acid was also detected in these experiments but could not be quantified because its very low volatility prevented a quantitative introduction in the chamber and hence FTIR calibration. In CESAM chamber experiments, no gas phase products could be positively identified using FTIR spectra. In EUPHORE chamber experiments, formic acid and carbon monoxide were detected using FTIR spectroscopy and molar yields were determined to be  $8 \pm 1$  % and  $18 \pm 1$  % respectively (Table 4). Moreover, 14 other reaction products were detected by on-line GC-MS and for most of them a molecular structure was proposed (Table 5). All species bearing a carbonyl function were also detected using DNPH cartridge sampling and off-line LC-MS analysis. Furthermore, in experiment Ind\_E1, where concentrations were much higher, formaldehyde, acetaldehyde and acetone were also identified using this technique.

**Table 4. Molar Yields of the Gas Phase Products Formed during Indene Ozonolysis**

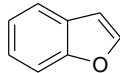
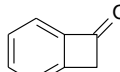
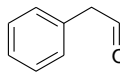
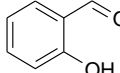
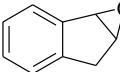
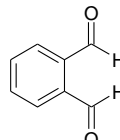
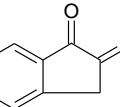
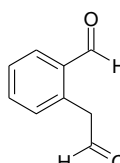
Experiments (#)	Gaseous products yields (%)			
	<i>Indene ozonolysis</i>	<i>Formic acid</i>	<i>Carbon monoxide</i>	<i>Phthalic acid</i>
Ind_E1		$7 \pm 1$	$18 \pm 1$	(a)
Ind_E2		$8 \pm 1$	$18 \pm 1$	(a)
Ind_E3		$9 \pm 1$	$18 \pm 1$	(a)
Average mean		$8 \pm 1$	$18 \pm 1$	-

(a) not detected.

It should be emphasized that two impurities were detected in indene using on-line GC-MS. This was not completely negligible since their concentrations could be estimated, based on the same response factor as indene, representing about 10% of indene concentration in total. These two compounds were assigned to an undetermined alkylated aromatic compound ( $C_{10}H_{14}$ , molecular ion  $m/z$  134) and to 2-methylstyrene. Indeed, the formation of 2-methylstyrene ozonolysis products could be measured during indene ozonolysis

1 experiments. o-Tolualdehyde (M1) was detected in Ind\_E1, Ind\_E2 and Ind\_E3 and two  
2  
3 particulate products of 2-methylstyrene ozonolysis, i.e. 2-methylphenylformate (M2) and 2-  
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5 methylphenol (M6), were also detected in Ind\_E1, probably due to higher initial  
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7 concentrations in this experiment. Moreover, it should be noted also that toluene was present  
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9 in the chamber during Ind\_E1 and Ind\_E3 experiments at very low concentrations.  
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**Table 5. Gas Phase Products Formed during Indene Ozonolysis Identified using On-Line GC-MS and/or Off-Line LC-MS**

Compound name <i>Detected in Exp#</i>	Molecular structure	MW (g mol <sup>-1</sup> )	Major ions observed in GC-MS (m/z)	M+179 ions observed in LC-MS (m/z)	Time profile pattern <sup>(d)</sup>
<b>Benzofuran</b> <i>Ind_E1, E2 and E3</i>		118	63, 89, 91, 118	n.d. <sup>(a)</sup>	I
<b>Benzocyclobuten-1(2H)-one</b> <i>Ind_E1, E2 and E3</i>		118	63, 89, 91, 118	297	I
<b>Phenylacetaldehyde</b> <i>Ind_E1, E2 and E3</i>		120	65, 91, 92, 120	299	I (+ II in Ind_E3)
<b>2-Hydroxybenzaldehyde</b> <i>Ind_E1, E2 and E3</i>		122	65, 76, 93, 121, 122	301	I (+ II in Ind_E3)
<b>Indene oxide</b> <i>Ind_E1, E2 and E3</i>		132	78, 103, 104, 132	n.d. <sup>(a)</sup>	I
<b>Phthalaldehyde</b> <i>Ind_E1, E2 and E3</i>		134	51, 77, 105, 134	313 <sup>(b)</sup> 493 <sup>(c)</sup>	I (+ II in Ind_E3)
<b>1H-indene-1,2(3H)dione</b> <i>Ind_E1</i>		146	63, 77, 89, 90, 105, 118, 146	325 <sup>(b)</sup> 505 <sup>(c)</sup>	I
<b>2-(formylmethyl)benzaldehyde</b> <i>Ind_E1</i>		148	90, 118, 148	327 <sup>(b)</sup> 507 <sup>(c)</sup>	II
<b>IndPox_1</b> <i>Ind_E1</i>	-	-	50, 51, 77, 105, 133, 134	313	II
<b>IndPox_2</b> <i>Ind_E1</i>	-	-	50, 51, 75, 76, 102, 129		II
<b>IndPox_3</b> <i>Ind_E1</i>	-	-	63, 89, 105, 117, 145, 146	n.d. <sup>(a)</sup>	I

<sup>(a)</sup> not detected in LC-MS analysis.

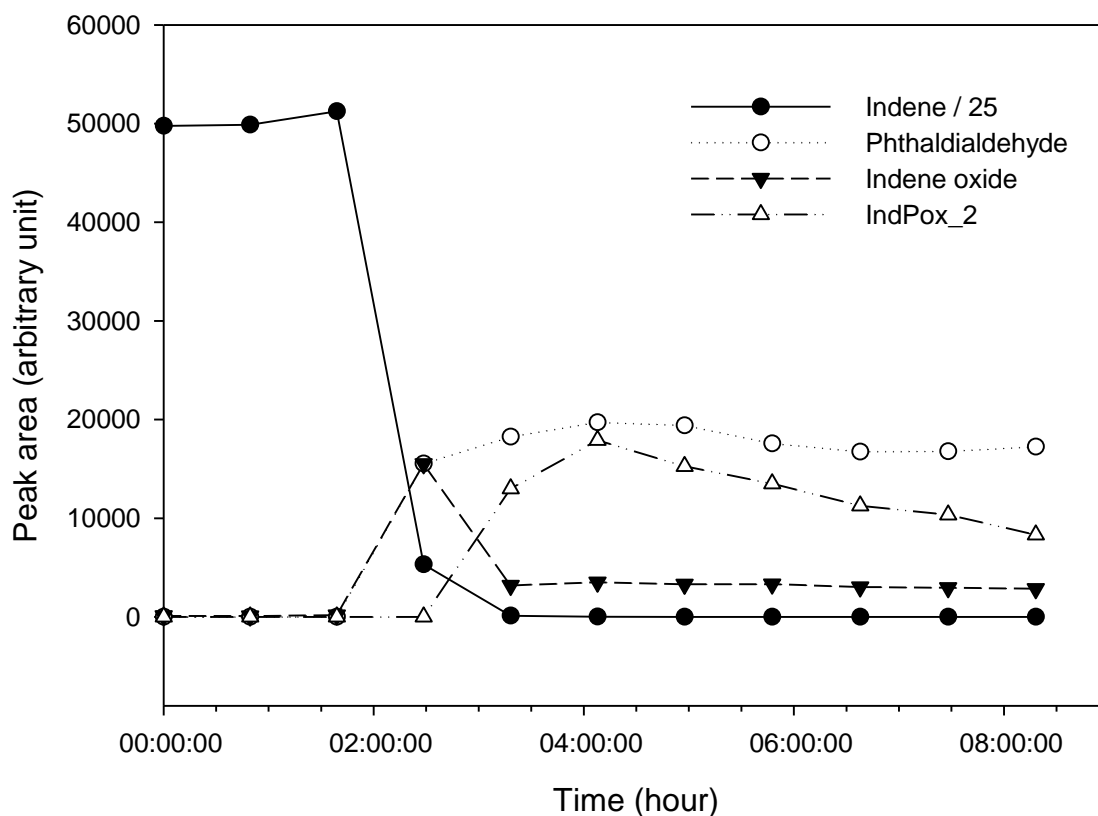
<sup>(b)</sup> molecular ion of monosubstituted DNPH derivative.

<sup>(c)</sup> molecular ion of disubstituted DNPH derivative.

<sup>(d)</sup> the time profile pattern was either primary-like (I) or secondary-like (II).

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2 The other 11 ozonolysis products detected in the gas phase during indene ozonolysis  
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4 experiments were all detected in Ind\_E1 and most of them were also present in Ind\_E2 and  
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6 Ind\_E3. Among them, indene oxide, phthaldialdehyde, 1H-indene-1,2(3H)dione and 2-  
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8 (formylmethyl)benzaldehyde were identified as indene ozonolysis products and formation  
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10 pathways were proposed. Furthermore, 2-hydroxy-benzaldehyde, benzofuran,  
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12 benzocyclobuten-1(2H)-one and phenylacetaldehyde were also identified but no reaction  
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14 pathway could be proposed to explain their formation from indene, indicating that they may  
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16 result from impurities oxidation. Finally, three other compounds were also detected for which  
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18 no molecular structure could be proposed. Nevertheless, they are presented in Table 5,  
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20 along with mass spectra information since some of them, such as IndPox\_2, were assumed  
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22 to be abundant, as suggested by their signal abundances.  
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24  
25 The time evolution of these products was also studied in order to retrieve information on their  
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27 formation as well as degradation pathways, although it is important to stress that  
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29 complementary experiments should be performed to confirm the following results. As noticed  
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31 in Table 5, most of the products show a typical primary product profile. This is illustrated on  
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33 Figure 1 by the GC-MS signal of indene and selected products (phthaldialdehyde, indene  
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35 oxide and IndPox\_2) versus time. Among the primary-like products, some were found to be  
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37 rapidly consumed after formation, such as indene oxide. Moreover, phthaldialdehyde, as well  
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39 as 2-hydroxybenzaldehyde and phenylacetaldehyde, have a primary-like product pattern in  
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41 Ind\_E1 and Ind\_E2 but a more likely secondary-like product profile in Ind\_E3, due to a  
42  
43 delayed formation or transfer from the particles back to the gaseous phase. Ind\_E2 and  
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45 Ind\_E3 were performed with comparable initial concentrations of ozone and indene, except  
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47 that the chamber temperature was higher during Ind\_E3 experiment. Thus, it is expected that  
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49 for these semi-volatile compounds, a temperature increase would induce an increase in the  
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51 vapor pressure and therefore a later gas-to particle partitioning. IndPox\_1, IndPox\_2 and 2-  
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53 (formylmethyl)benzaldehyde show typical profiles of secondary-like products. These three  
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55 products were detected in Ind\_E1 only, where reaction has progressed further considering  
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57 higher initial concentrations, confirming that they could indeed be produced as secondary  
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59 products.  
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**Figure 1.** Typical time profiles of indene and gas phase products during indene ozonolysis (experiment Ind\_E1). Initial concentrations were 756 ppbv and a large excess (>1000 ppbv) for indene and ozone, respectively. Temperature was  $294 \pm 1$  K.

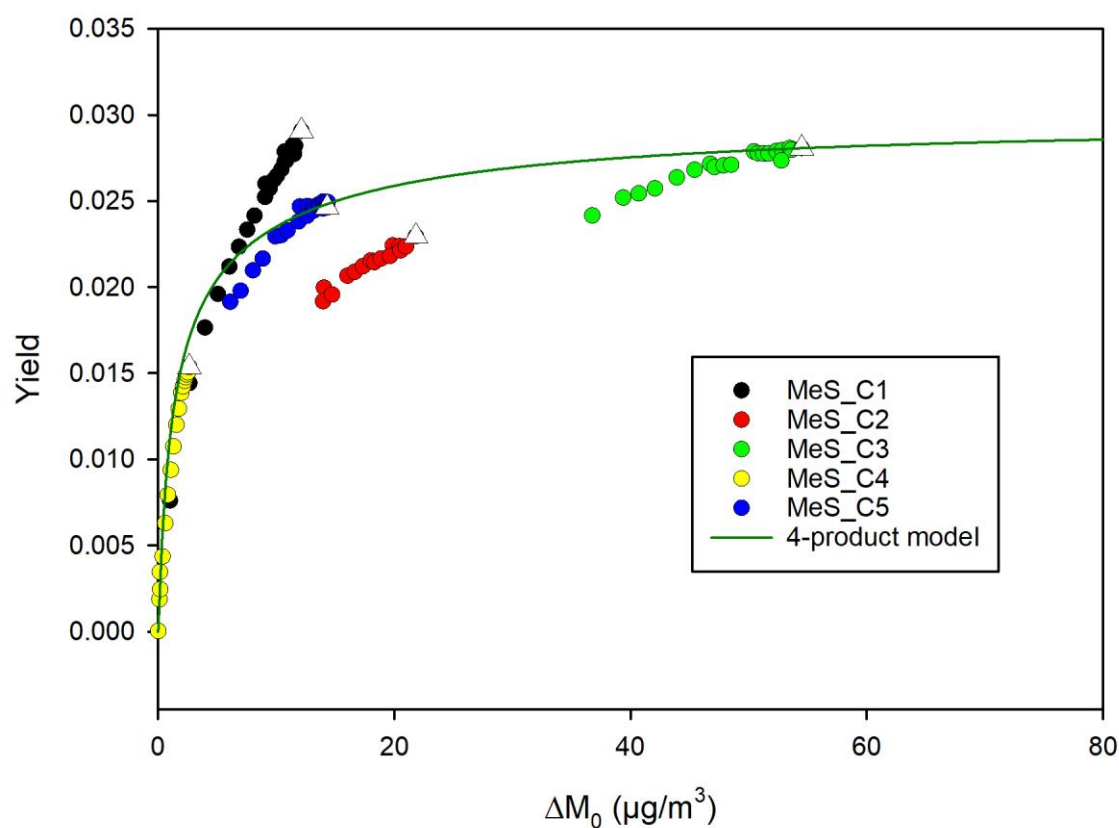
Finally, it is important to note that no carboxylic acid was detected in the gaseous phase, although they accounted for a large fraction of the identified compounds in the particulate phase (see section 3.3.). As a matter of fact, oven temperature was raised up to  $200$  °C, which is, for example, not enough for *o*-toluic acid elution and detection ( $T_{\text{boiling point}} \sim 260$  °C).

### 3.2. Aerosol growth and yield

#### 2-Methylstyrene ozonolysis

For all experiments, SOA formation began as soon as 2-methylstyrene ozonolysis started. When plotting aerosol mass concentration as a function of 2-methylstyrene reacted (Figure S1), results from the different experiments overlap reasonably well with each other,

1 highlighting a fair consistency between experiments. For all experiments, a slight slope  
 2 increase was observed in the first minutes of the reactions (Figure S1) which could either be  
 3  
 4 increase was observed in the first minutes of the reactions (Figure S1) which could either be  
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 6 due to (i) gas-phase reaction kinetics limiting saturation or (ii) gas-to-particle transfer limiting  
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 8 condensation or (iii) possible wall saturation. Then, aerosol growth is almost linear with 2-  
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 10 methylstyrene consumption and reaches its maximum when 2-methylstyrene is consumed  
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 12 and remains constant thereafter (Figure S1). This behavior is similar to the one observed by  
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 14 Ng et al. for  $\alpha$ -pinene during ozonolysis as well as other terpenoids bearing one double-  
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 16 bond.<sup>18</sup> From this growth pattern, they assume that the first oxidation product formation is  
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 18 most probably the rate-determining step in SOA formation.  
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 50 **Figure 2.** SOA yield curve for 2-methylstyrene ozonolysis. Experimental data are  
 51 represented with coloured circles and for each experiment, the white triangle represent the  
 52 final yield. The four-product model fitting is represented by the green curve and the fitting  
 53 parameters are ( $C_1^* = 1$  ;  $\alpha_1 = 2 \times 10^{-2}$  ;  $C_2^* = 10$  ;  $\alpha_2 = 9 \times 10^{-3}$  ;  $C_3^* = 10^2$  ;  $\alpha_3 = 6 \times 10^{-9}$  ;  $C_4^* = 10^3$  ;  
 54  $\alpha_4 = 2 \times 10^{-8}$ ), where  $C_i^*$  are the saturation concentrations ( $\mu\text{g m}^{-3}$ ) and  $\alpha_i$  is the stoichiometric  
 55 coefficients.  
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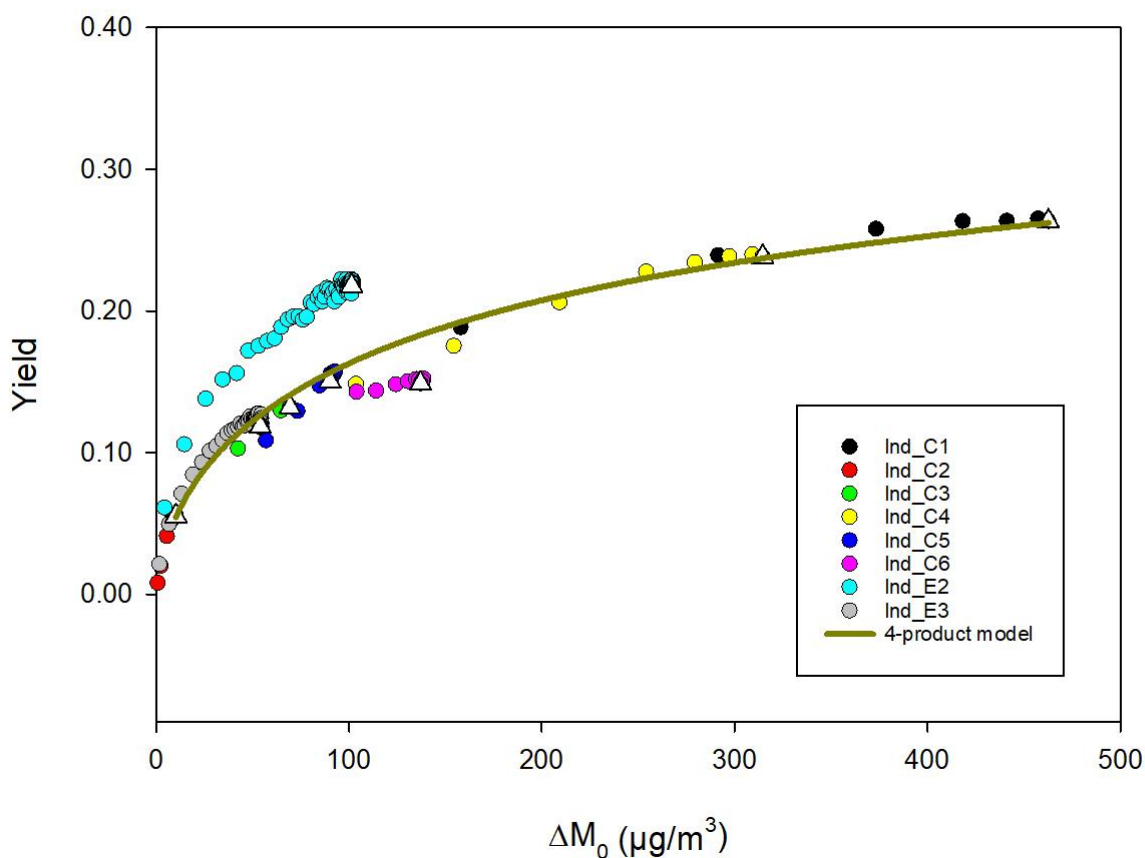
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4 SOA yields are determined for the 5 experiments performed in CESAM and range from  $1.5 \pm$   
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6  $0.2 \%$  to  $2.8 \pm 0.1 \%$ . They are plotted versus aerosol mass concentration  $M_0$  in Figure 2,  
7  
8 where they are fitted using the four-product model. The best correlation coefficient was  
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10 obtained for saturation concentrations ranging from 1 to  $1000 \mu\text{g m}^{-3}$ . Despite a significant  
11  
12 dispersion of experimental data points, the fit shows acceptable correlation coefficients  
13  
14 ( $R^2=0.69$ ). The main products (with highest gas-phase stoichiometric fractions) are those of  
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16 lowest volatilities, i.e. hypothetical products with saturation concentrations of 1 and  $10 \mu\text{g m}^{-3}$ .  
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18 2-Methylstyrene ozonolysis SOA yield curve is further discussed as regards indene  
19  
20 ozonolysis SOA yield curve (section 4.3.).  
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### 25 *Indene ozonolysis*

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27 Although the introduction order of ozone and indene in the chamber was not always the  
28  
29 same for the set of indene ozonolysis experiments, the results are consistent. For all  
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31 experiments, SOA formation was observed as soon as indene started to react with ozone. As  
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33 illustrated in Figure S2, in all experiments except Ind\_E1, aerosol mass concentration  
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35 increased almost linearly after a short delay and reached a maximum when indene was  
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37 consumed. It has to be emphasized here that the experimental data generated in EUPHORE  
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39 and CESAM facilities, using slightly different experimental procedures overlap quite well and  
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41 the time-dependent growth curve pattern is similar to that of 2-methyl-styrene. In experiment  
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43 Ind\_E1 only, the growth curve pattern is very different: SOA yield is much larger than for all  
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45 other experiments, due to much larger precursor initial concentrations and SOA mass  
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47 continues to grow even after indene is consumed. This pattern is similar to the ones of  
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49 hydrocarbon precursors bearing more than one double bond (or reactive site) as shown by  
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51 Ng et al.<sup>18</sup> It would suggest that further oxidation of first-generation products may occur and  
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53 contribute to SOA growth. Indeed, in this particular experiment, indene and ozone were both  
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55 introduced at high concentrations ( $756 \text{ ppbv}$  and  $> 1 \text{ ppmv}$  for indene and ozone,  
56  
57 respectively) and ozone was introduced slowly (during about 1 hour) in order to decrease  
58  
59 indene ozonolysis reaction rate. Ozone was also present in a large excess during the whole  
60

1 experiment, and OH radicals, potentially generated from ozonolysis, may participate and  
2 further oxidise first generation products and hence contribute significantly to SOA growth. It  
3  
4 can even be proposed that the aromatic ring would be attacked at this stage of the reaction.  
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6 The hypothesis of a further oxidation may be corroborated by the results presented in the  
7  
8 previous section where secondary-like products such as IndPox\_2 and 2-  
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10 (formylmethyl)benzaldehyde were detected for Ind\_E1 preferentially. Nevertheless, this  
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12 “secondary” chemistry may also contribute to SOA mass in Ind\_E2 and Ind\_E3, but to a  
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14 much lesser extent, as concentrations were lower and ozone was present in a smaller  
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16 excess. Indeed, SOA yields were higher for Ind\_E2 than for experiments performed in  
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18 CESAM, where ozone was not in excess. For Ind\_E3 however, SOA yields were similar to  
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20 the ones measured in CESAM experiments, due to higher temperature (304 K) in this  
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22 particular experiment.  
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27 SOA yields were determined for the 9 experiments performed in EUPHORE and CESAM and  
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29 values range from  $8.8 \pm 0.8 \%$  to  $27.4 \pm 0.3 \%$ . It is not expected that the presence of  
30  
31 impurities will impact significantly these values since (i) ozonolysis of 2-methylstyrene  
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33 induces much lower SOA mass yields by an order of magnitude, (ii) toluene or other  
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35 alkylated monoaromatic compounds are not very reactive with ozone, and (iii) SOA  
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37 generated by OH-initiated oxidation of these impurities would represent only a minor fraction  
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39 of the total SOA, considering similar SOA yields for these precursors but much lower initial  
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41 concentrations (i.e. SOA yields values of about 30% and 36% (with SOA density of  $1.4 \text{ g cm}^{-3}$ )  
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43 were reported for OH-initiated oxidation of toluene and m-xylene respectively, in dry and  
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45 low NO<sub>x</sub> conditions by Ng et al.<sup>37</sup> Yields were plotted versus aerosol mass concentration  $M_0$   
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47 (Figure 3) and again, experimental data points were fitted using the four-product model for all  
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49 experiments except Ind\_E1 due to inconsistent experimental conditions as mentioned above  
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51 and in Table 2. The corresponding coefficient parameters are listed in Figure 3. The fit shows  
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53 good correlation coefficients ( $R^2=0.87$ ). The best correlation coefficient was obtained for  
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55 saturation concentrations ranging from 1 to  $1000 \mu\text{g m}^{-3}$ . The most abundant product is the  
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57 less volatile with a saturation concentration of  $1 \mu\text{g m}^{-3}$ .  
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**Figure 3.** SOA yield curve for indene ozonolysis. Experimental data are represented with coloured circles and for each experiment, the white triangle represent the final yield. The four-product model fitting is represented by the green curve. The fitting parameters are ( $C_1^* = 1$  ;  $\alpha_1 = 0.87$  ;  $C_2^* = 10$  ;  $\alpha_2 = 0.058$  ;  $C_3^* = 10^2$  ;  $\alpha_3 = 0.17$  ;  $C_4^* = 10^3$  ;  $\alpha_4 = 0.18$ ) where  $C_i^*$  are the saturation concentrations ( $\mu\text{g m}^{-3}$ ) and  $\alpha_i$  the stoichiometric coefficients.

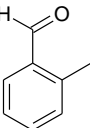
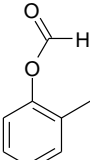
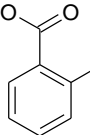
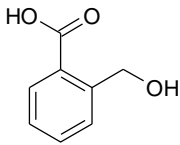
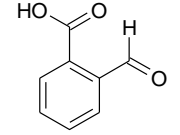
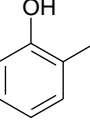
### 3.3. Aerosol chemical composition

#### 2-Methylstyrene ozonolysis

SFE-GC-MS analysis allowed to detect more than ten compounds in SOA and a molecular structure was proposed for 6 of them (Table 6) : o-tolualdehyde (M1), 2-methylphenylformate (M2), o-toluic acid (M3), 2-(hydroxymethyl)benzoic acid (M4), 2-methylbenzoic acid (M5) and 2-methylphenol (M6). The molecular structures could be confirmed using standards for o-tolualdehyde and o-toluic acid only. On-line analysis by TD-API-AMS allowed the detection of

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2 m/z 135 and 121 ions in the negative and the positive ionisation modes, respectively, hence  
3  
4 assigned to o-toluic acid and o-tolualdehyde. Two other ions (m/z 135 and 141) were also  
5  
6 detected in the positive ionisation mode but no molecular structure could be attributed to  
7  
8 these ions. A list of these compounds is presented in Table 6, along with the characteristic  
9  
10 fragments ( $[M+H]^+$ ,  $[M-CH_3]^+$ ,  $[M-OSi(CH_3)_3]^+$ ) used for SFE-GC-MS identification, the  
11  
12 molecular ion m/z ratios used for TD-API-AMS detection and the corresponding standard or  
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14 surrogate standards selected for quantification.  
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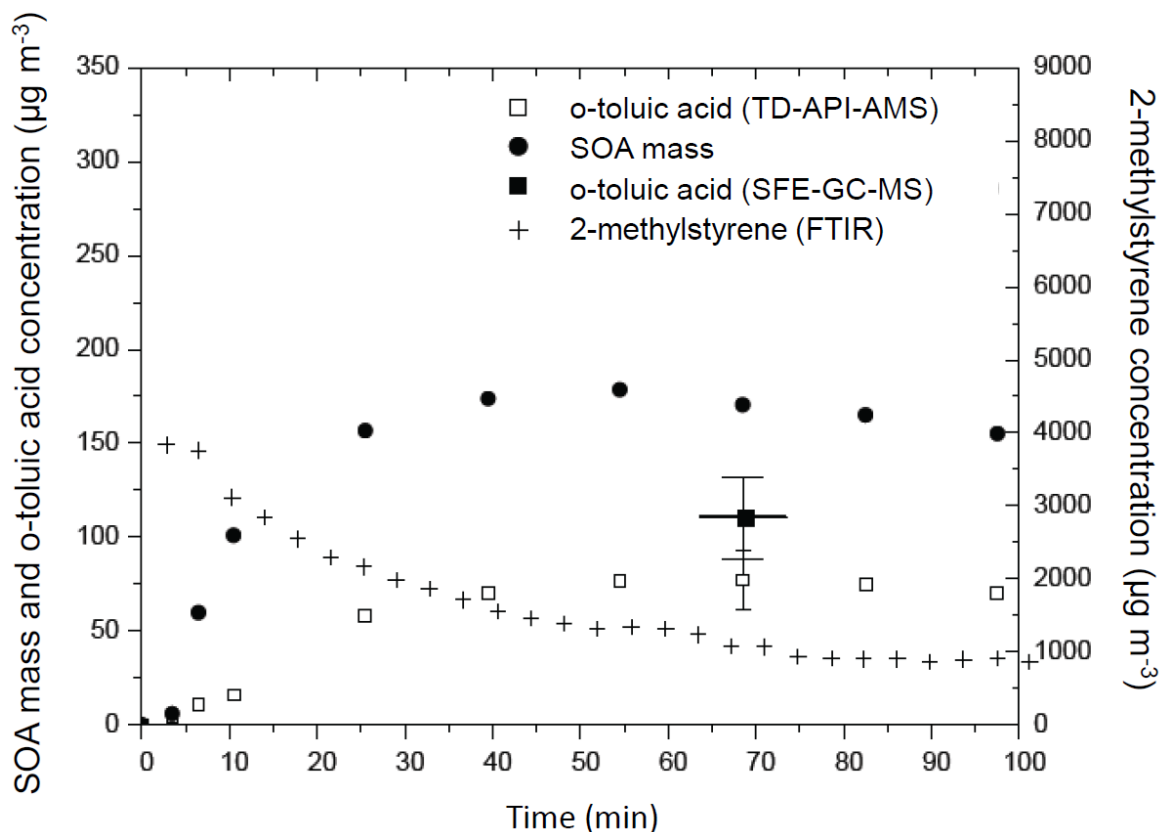
**Table 6.** List of the compounds identified in the particulate phase during 2-methylstyrene ozonolysis using off-line SFE-GC-MS and/or on-line TD-API-AMS.

Compound name	Molecular structure	Formula MW <i>TMS_MW</i> <sup>(1)</sup> ( <i>g mol</i> <sup>-1</sup> )	SFE-GC-MS <sup>(2)</sup> CI fragments ( <i>m/z</i> )	TD-API-AMS <sup>(2)</sup>		Quantification standard	Concentration ( $\mu\text{g m}^{-3}$ )	Molar yield (%)
				Positive mode ( <i>m/z</i> )	Negative mode ( <i>m/z</i> )			
o-tolualdehyde (M1)		C <sub>8</sub> H <sub>8</sub> O 120 -	M+1=121	M+1=121		o-tolualdehyde	2.7 ± 0.8 <sup>(2)</sup>	0.06 ± 0.03 <sup>(2)</sup>
2-methylphenyl formate (M2)		C <sub>8</sub> H <sub>8</sub> O <sub>2</sub> 136 -	M+1=137			phenyl formate	1.5 ± 0.4 <sup>(2)</sup>	0.03 ± 0.02 <sup>(2)</sup>
o-toluic acid (M3)		C <sub>8</sub> H <sub>8</sub> O <sub>2</sub> 136 208	M+1=209 M-15=193 M-89=119		M-1=135	benzoic acid	140 ± 50 <sup>(2)</sup> 80 ± 16 <sup>(3)</sup>	2.4 ± 1.4 <sup>(2)</sup> 3.4 ± 1.1 <sup>(3)</sup> 2.1 ± 0.7 <sup>(4)</sup>
2-(hydroxymethyl) benzoic acid (M4)		C <sub>8</sub> H <sub>8</sub> O <sub>3</sub> 152 296	M-15=281 M-89=207			3-hydroxy benzoic acid	0.40 ± 0.14 <sup>(2)</sup>	0.006 ± 0.003 <sup>(2)</sup>
2-formyl benzoic acid (M5)		C <sub>8</sub> H <sub>6</sub> O <sub>3</sub> 150 222	M+1=223 M-15=207 M-89=133			2-acetyl benzoic acid	0.9 ± 0.3 <sup>(2)</sup>	0.06 ± 0.03 <sup>(2)</sup>
2-methylphenol (M6)		C <sub>7</sub> H <sub>8</sub> O 108 180	M-15=165 M-89=91			3- methylphenol	1.6 ± 0.4 <sup>(2)</sup>	0.04 ± 0.02 <sup>(2)</sup>

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- (1) TMS\_MW : Trimethylsilyl derivative molecular weight.
- (2) Results obtained by SFE-GC-MS from experiment MeS\_I1.
- (3) Results obtained by TD-API-AMS from experiment MeS\_I1.
- (4) Results obtained by TD-API-AMS from experiment MeS\_I2.

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2 The detailed quantification of the particulate phase was performed for experiment MeS\_I1  
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4 and o-toluic acid was found to be the main product identified in SOA formed by 2-  
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6 methylstyrene ozonolysis (Table 6). O-Toluic acid concentrations, determined independently  
7  
8 by SFE-GC-MS ( $110 \pm 22 \mu\text{g m}^{-3}$ ) and TD-API-AMS ( $80 \pm 16 \mu\text{g m}^{-3}$ ), are in good agreement  
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10 within uncertainties. This monocarboxylic acid represents  $\sim 95\%$  of the identified aerosol  
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12 fraction and  $\sim 3\%$  of the reacted 2-methylstyrene. All the other compounds were present in  
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14 SOA at much lower concentrations ranging from  $0.40 \pm 0.14 \mu\text{g m}^{-3}$  for 2-(hydroxymethyl)  
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16 benzoic acid to  $2.7 \pm 0.4 \mu\text{g m}^{-3}$  for o-tolualdehyde. Finally, considering that the aerosol mass  
17  
18 concentration measured by SMPS was under-estimated since SOA size distribution was  
19  
20 larger than the SMPS monitored mass range, only an upper limit of  $54 \pm 17\%$  could be  
21  
22 estimated for the identified aerosol fraction. The missing particulate compounds were  
23  
24 assumed to be either low volatility compounds such as oligomers, which could not be  
25  
26 detected using our analytical techniques; or semi-volatile compounds such as dicarbonyl  
27  
28 species produced from aromatic ring cleavage, which are eluted in GC-MS analysis in the  
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30 same time as BSTFA and thus can also not be detected. Finally, it cannot be excluded that  
31  
32 the missing particulate compounds are analysable with these techniques but present at very  
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34 low concentrations, under the detection limits.  
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**Figure 4.** Time profiles of the quantified reaction products formed during 2-methylstyrene ozonolysis for experiment MeS\_I1. Initial concentrations were 900 ppbv and 870 ppbv for 2-methylstyrene and ozone, respectively. Temperature was  $298 \pm 2$  K (see more experimental conditions in Table 2). The concentration of 2-methylstyrene was monitored in the gas phase by FTIR whereas those of o-toluic acid were measured in the particulate phase by TD-API-MS and SFE-GC-MS.

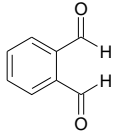
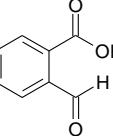
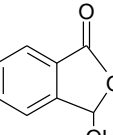
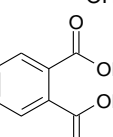
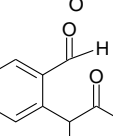
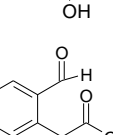
The time evolution of o-toluic acid concentrations in particles obtained from TD-API-AMS data is plotted in Figure 4 and shows that this product is present from the beginning of SOA formation and its concentration increases with SOA mass. This suggests that o-toluic acid contributes to particulate matter formation, either being formed in the gaseous phase and rapidly condensed on particles afterwards or being formed in the particulate phase rapidly after the gas-to-particle transfer of its precursor, possibly o-tolualdehyde.

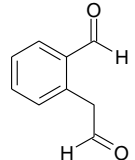
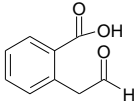
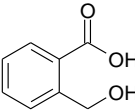
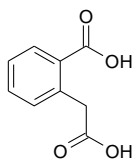
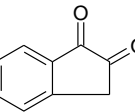
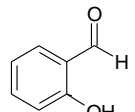
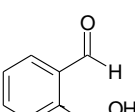


### *Indene ozonolysis*

SFE-GC-MS analysis performed on filters sampled from experiment Ind\_I1 allowed the detection of about 30 compounds, illustrating that the chemical composition of SOA formed *via* indene ozonolysis was much more complex than the one formed *via* 2-methylstyrene ozonolysis. Among these compounds, 10 could be identified and are listed in Table 7: phthalaldehyde (I1), 2-formyl benzoic acid (I2), 3-hydroxy-1(3H)-isobenzofuranone (I3), phthalic acid (I4), 2-hydroxy-2-(2-formylphenyl) acetaldehyde (I5), (2-formylphenyl) acetic acid (I6), 2-(2-oxo-ethyl) benzoic acid (I8), 2-(hydroxymethyl) benzoic acid (I9), 2-carboxyphenyl acetic acid (I10) and 1H-indene-1,2(3H)dione (I11). It is important to note that two chromatographic peaks were detected for I6 and I8 but they could not be assigned from their mass spectrum to one or the other molecular structure. All the molecular ions corresponding to these compounds were also detected by TD-API-AMS and monitored throughout the experiments performed in ICARE and EUPHORE chambers, either in positive ( $[M+H]^+$ ) or negative ( $[M-H]^-$ ) ionisation mode. Moreover, the detection of phthalaldehyde (I1), 2-formyl benzoic acid (I2) and phthalic acid (I4) was positively confirmed in SOA particles generated in EUPHORE experiments, by mass spectrometry using MS<sup>2</sup> mode. The fragment profiles in SOA of the so-assigned ions (i.e. m/z 135 in the positive mode, 149 m/z and 165 m/z in the negative mode) were compared and found to agree with those of their corresponding standards. As mentioned in section 3.1., phthalaldehyde (I1) and 1H-indene-1,2(3H)dione (I11) were also detected in the gaseous phase in EUPHORE experiments by on-line GC-MS. Three additional species (ions m/z 149<sup>+</sup>, 121<sup>-</sup> and 135<sup>-</sup>) were identified using TD-API-AMS. Although no standards were available for confirmation, 2-(formylmethyl)benzaldehyde (I7), 2-hydroxybenzaldehyde (I12) and 2-(hydroxymethyl)benzaldehyde (I13) molecular structures were proposed for m/z 149<sup>+</sup>, 121<sup>-</sup> and 135<sup>-</sup> on the basis of their MS<sup>2</sup> spectra. All the products identified in indene ozonolysis experiments are thus oxidized, highly polar and bi-functionalized (Table 7).

**Table 7.** List of compounds identified in the particulate phase during indene ozonolysis using off-line SFE-GC-MS and/or on-line TD-API-AMS.

Compound name	Molecular structure	Formula MW <i>TMS_MW</i> <sup>(1)</sup> ( <i>g mol</i> <sup>-1</sup> )	SFE-GC-MS <sup>(2)</sup> CI fragments ( <i>m/z</i> )	TD-API-AMS <sup>(2,3)</sup>		Quantification standard	Concentration <sup>(2)</sup> ( $\mu\text{g m}^{-3}$ )	Molar yield <sup>(2)</sup> (%)
				Positive mode ( <i>m/z</i> )	Negative mode ( <i>m/z</i> )			
phthal-dialdehyde (I1)		C <sub>8</sub> H <sub>6</sub> O <sub>2</sub> 134 -	M+1=135	M+1=135		benzaldehyde	50 ± 16	1.7 ± 0.6
2-formyl benzoic acid (I2)		C <sub>8</sub> H <sub>6</sub> O <sub>3</sub> 150 222	M+1=223 M-15=207 M-89=133	M+1=151	M-1=149	2-acetyl benzoic acid	10.7 ± 4.0	0.3 ± 0.1
3-hydroxy-1(3H)iso benzofuranone (I3)		C <sub>8</sub> H <sub>6</sub> O <sub>3</sub> 150 222	M+1=223 M-15=207 M-89=133	M+1=151		3-hydroxy-benz- aldehyde	2.0 ± 0.7	0.06 ± 0.03
phthalic acid (I4)		C <sub>8</sub> H <sub>6</sub> O <sub>4</sub> 166 310	M+1=311 M-15=295 M-89=207		M-1=165	benzoic acid	2.5 ± 1.0	0.07 ± 0.03
2-hydroxy-2-(2- formylphenyl) acetaldehyde (I5)		C <sub>9</sub> H <sub>8</sub> O <sub>3</sub> 164 236	M+1=237 M-15=221 M-89=147	M+1=165		3-hydroxy-benz- aldehyde	0.8 ± 0.3	0.02 ± 0.01
(2-formylphenyl) acetic acid (I6)		C <sub>9</sub> H <sub>8</sub> O <sub>3</sub> 164 236	M+1=237 M-15=221 M-89=147	M+1=165	M-1=163	2-acetyl benzoic acid	3.2 ± 1.4 (I6 or I8)	0.07 ± 0.04 (I6 or I8)

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2									
3									
4	2-(formylmethyl)		C <sub>9</sub> H <sub>8</sub> O <sub>2</sub> 148 -	-	M+1=149 MS <sup>2</sup> (149): 77, 91, 93, 103, 121, 131	-			
5	benzaldehyde								
6	(I7)								
7									
8									
9	2-(2-oxo-ethyl)		C <sub>9</sub> H <sub>8</sub> O <sub>3</sub> 164 236	M+1=237 M-15=221 M-89=147	M+1=165	M-1=163	2-acetyl benzoic acid	0.31 ± 0.12 (I6 or I8)	0.009 ± 0.004 (I6 or I8)
10	benzoic acid								
11	(I8)								
12									
13									
14	2-(hydroxymethyl)		C <sub>8</sub> H <sub>8</sub> O <sub>3</sub> 152 296	M+1=297 M-15=281 M-89=207			3-hydroxy- benzoic acid	3.2 ± 0.8	0.07 ± 0.04
15	benzoic acid								
16	(I9)								
17									
18									
19	2-carboxyphenyl		C <sub>9</sub> H <sub>8</sub> O <sub>4</sub> 180 324	M+1 = 325 M-15 = 309 M-89 = 235		M-1=179	benzoic acid	2.0 ± 0.7	0.05 ± 0.02
20	acetic acid								
21	(I10)								
22									
23									
24	1H-indene-		C <sub>9</sub> H <sub>6</sub> O <sub>2</sub> 146 -	M+1=147	M+1=147		sabinaketone	1.2 ± 0.3	0.04 ± 0.01
25	1,2(3H)dione								
26	(I11)								
27									
28	2-hydroxy		C <sub>9</sub> H <sub>8</sub> O <sub>2</sub> 122 -						
29	benzaldehyde								
30	(I12)								
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32									
33	2-(hydroxymethyl)		C <sub>9</sub> H <sub>8</sub> O <sub>2</sub> 136 -						
34	benzaldehyde								
35	(I13)								
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(1) TMS\_MW: Trimethylsilyl derivative molecular weight.

(2) Results obtained from samples produced in the ICARE chamber (experiment Ind\_I1).

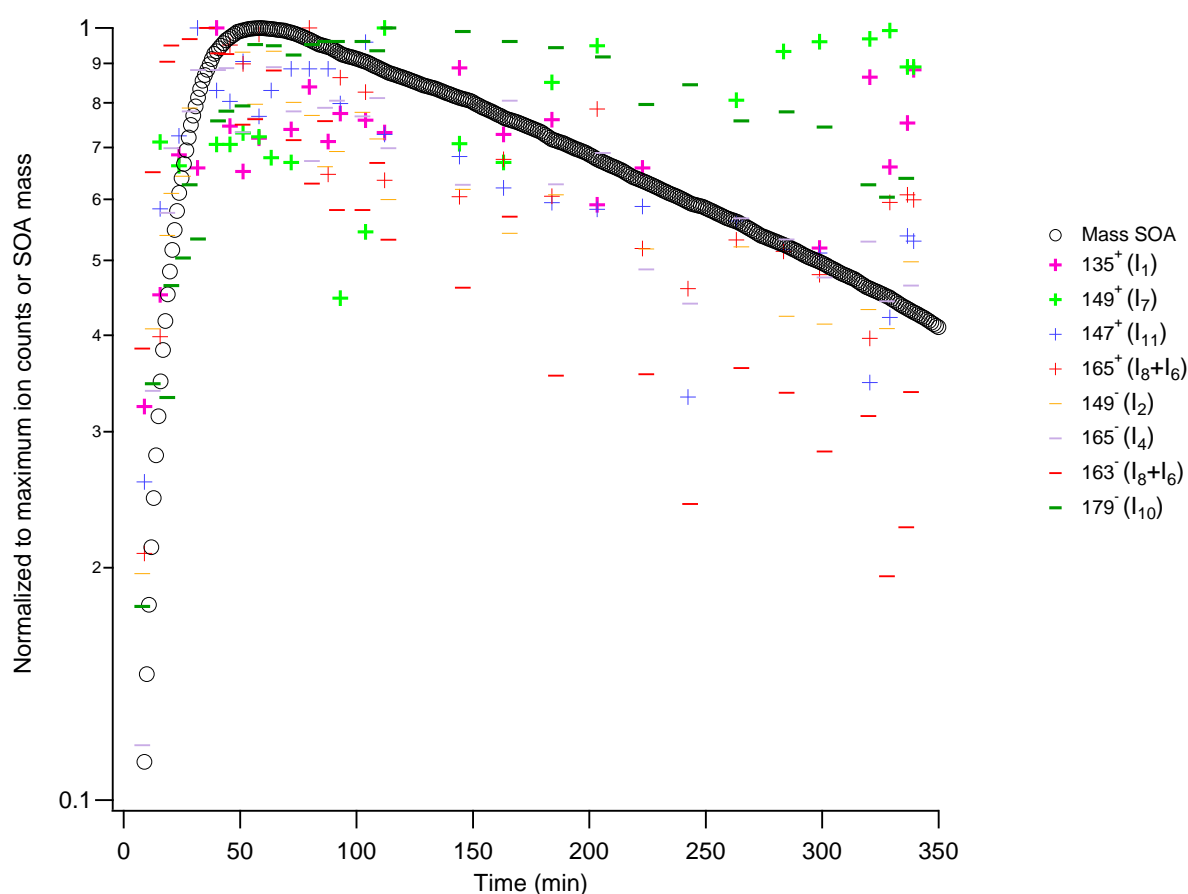
(3) Results obtained from samples produced in the EUPHORE chamber (experiments Ind\_E1 and Ind\_E3).

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4 It should be stressed that these results show again a strong consistency between the  
5  
6 information obtained from the different chambers on the chemical composition of both the  
7  
8 gaseous and the particulate phases.  
9

10 The detailed quantification of particulate compounds is provided in Table 7 for experiment  
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12 Ind\_I1 by SFE-GC-MS analysis. Phthaldialdehyde (I1) is the most abundant product ( $50 \pm 16$   
13  
14  $\mu\text{g m}^{-3}$ ) and represents  $66 \pm 20$  % of the identified aerosol mass concentration. Its molar yield  
15  
16 is  $1.7 \pm 0.6$  %, showing that the carbon balance is far from being complete in this study,  
17  
18 partly due to the lack of quantification of gaseous compounds. This may be explained by the  
19  
20 formation of a wide range of compounds present at very low concentrations. For this  
21  
22 experiment, performed in the ICARE chamber, we estimate that the upper limit of the  
23  
24 quantified aerosol mass concentration is  $25 \pm 6$  % of the total aerosol mass concentration.  
25  
26 This value is lower than the one obtained for SOA formed *via* 2-methylstyrene ozonolysis.  
27  
28 Again, the missing fraction of the aerosol may be partly composed of a wide range of low-  
29  
30 concentration species, possibly formed by aromatic ring cleavage or aggregation. During the  
31  
32 experiments performed in the EUPHORE chamber, the most abundant ions detected using  
33  
34 TD-API-AMS are  $m/z$  135+,  $m/z$  149- and  $m/z$  165-, respectively assigned to  
35  
36 phthaldialdehyde (I1), 2-formylbenzoic acid (I2) and phthalic acid (I4). The sum of their  
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38 concentrations reaches 45% and 96% of the total aerosol mass concentration during Ind\_E1  
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40 and Ind\_E2 experiments, respectively.  
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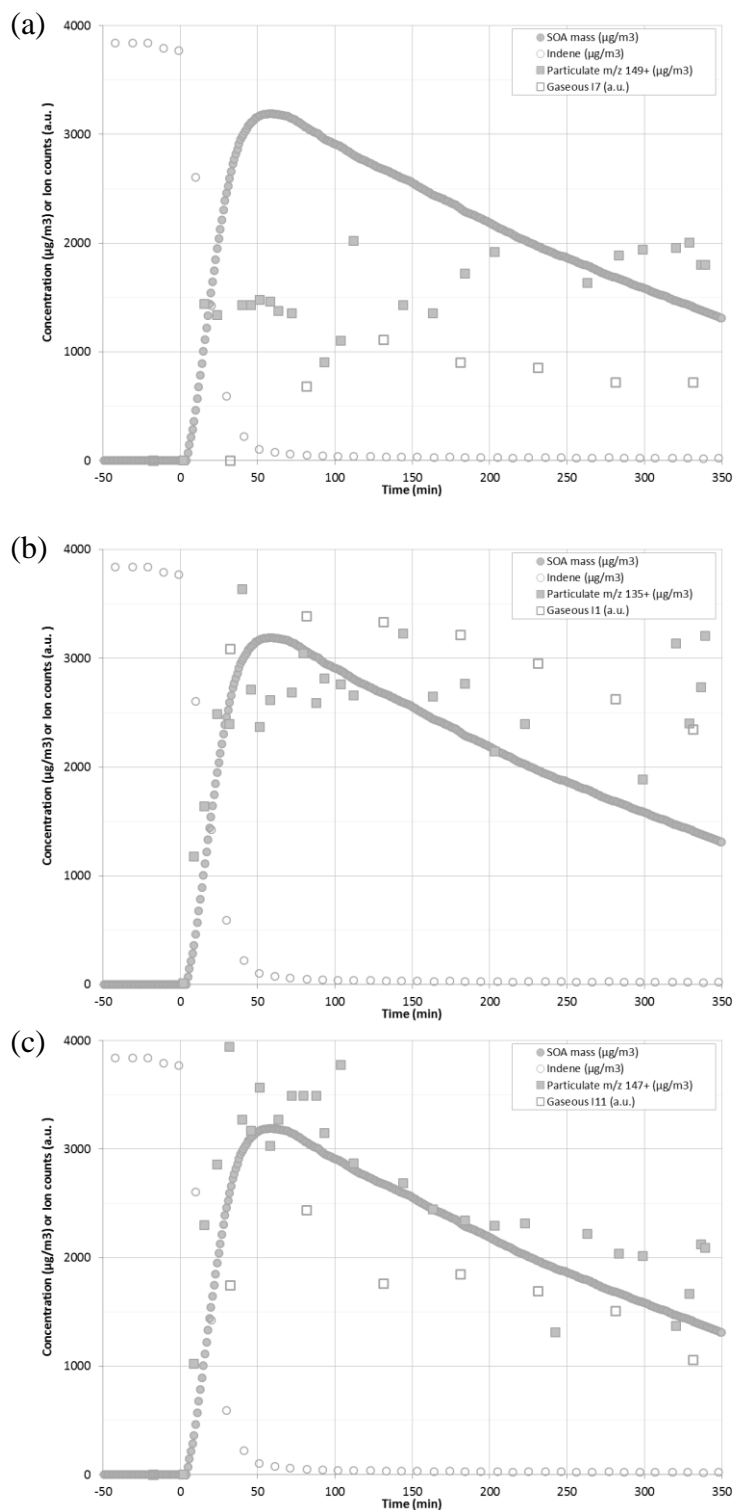
44 The time profiles of the molecular ions provided by TD-API-AMS along the experiments allow  
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46 relevant information on the mechanism to be retrieved. Most of the detected ions both in the  
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48 positive and the negative modes show a maximum signal simultaneously to the total aerosol  
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50 mass concentration maximum in experiments performed both in the ICARE and the  
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52 EUPHORE chambers as shown on Figures 5 and S3. This suggests that these compounds  
53  
54 may be formed rapidly as primary products or through reactions are not rate-limiting and that  
55  
56 they are stable and in equilibrium between the gaseous and particulate phases along  
57  
58 the aerosol formation. Ion signals  $m/z$  135+ (I1) and  $m/z$  149+ (I7) seem however to  
59  
60 decrease slower than the SOA mass (Figures 5 and S3) illustrating some enrichment of the  
particulate

1  
2 phase in these two dialdehydes. Furthermore, ions  $m/z$  165<sup>-</sup> (I<sub>4</sub>) and  $m/z$  179<sup>-</sup> (I<sub>10</sub>) show a  
3  
4 somewhat delayed maximum in experiment Ind\_E3 (Figure S3) and Ind\_E1 (Figure 5)  
5  
6 respectively, illustrating a possible secondary origin, either in the gaseous or in the  
7  
8 particulate phase or at least a slower formation. The slightly different time evolution patterns  
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10 between experiments suggests that some compounds may have several isomers with/or  
11  
12 different formation pathways with some of them being favored depending on the  
13  
14 experimental conditions (temperature, precursor concentration...).



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48 **Figure 5.** Normalised (to maximum value) positive and negative ion signals and SOA mass  
49 concentration during indene ozonolysis (Ind\_E1) obtained by TD-API-AMS. The ions  
50 corresponding to I<sub>1</sub>, I<sub>7</sub> and I<sub>10</sub> (**bold**) decrease significantly slower than that of SOA mass  
51 concentration.  
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1  
2 The gas-particle partitioning of some carbonyls was investigated and is presented for Ind\_E1  
3  
4 in Figure 6. I1 and I11 concentrations in both gaseous and particulate phases follow a similar  
5  
6 time profile, i.e. a rapid formation and a slow decrease, concomitantly with SOA mass  
7  
8 concentration. This may evidence an equilibrium between both phases for these two  
9  
10 products. For I7 however, its gaseous concentration was very low at the beginning of the  
11  
12 reaction and reached a maximum later on. Note that I7 is expected to be a major primary  
13  
14 product of indene ozonolysis even if a potential source of I7 (e.g. chamber walls, particle  
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16 matter, a specific reaction...) seemed to release it along the experiment.  
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**Figure 6.** Gas-particle partitioning of some indene ozonolysis products during Ind\_E1 : (a) 2-(formylmethyl) benzaldehyde (I7), (b) phthalaldehyde (I1) and (c) 1H-indene-1,2(3H)dione (I11).

1  
2 Figures 5 and 6 highlight that for some ions, signal-to-noise ratios are quite low, leading to  
3  
4 difficult interpretation and possibly due to the simultaneous presence of distinct products  
5  
6 yielding the same ions (either as molecular ion or fragment). Consequently, the interpretation  
7  
8 proposed above needs to be considered with caution.  
9

## 10 11 12 **4. Discussion**

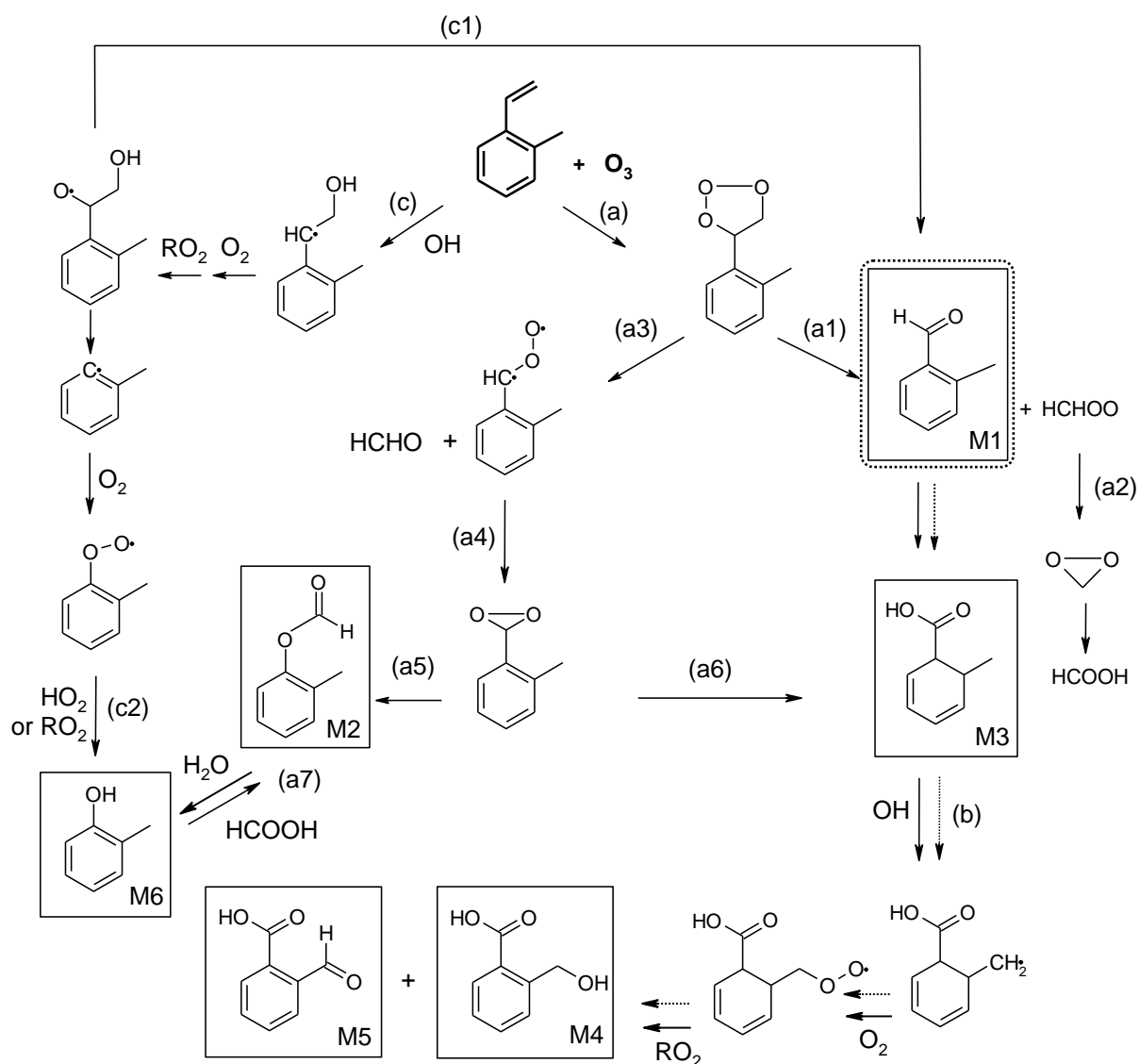
### 13 14 *4.1. Proposed mechanism for 2-methylstyrene ozonolysis*

15  
16 Figure 7 depicts the proposed routes to explain the formation of the identified products of 2-  
17  
18 methylstyrene ozonolysis. The mechanism presented here is inspired from those proposed  
19  
20 by Tuazon et al.,<sup>35</sup> Winterhalter et al.,<sup>38,39</sup> and Jenkin.<sup>40</sup> o-Tolualdehyde (M1), is a primary  
21  
22 gas phase reaction product, it is assumed to result from the decomposition of the ozonide  
23  
24 (reaction a1) formed from the addition of ozone on the 2-methylstyrene vinylic double bond.  
25  
26 Simultaneously, a small vibrationally excited Criegee radical is formed leading to the  
27  
28 formation of formic acid as a secondary product of the reaction (reaction a2). Formaldehyde,  
29  
30 also detected as a primary product in the gas phase, is formed following a similar process,  
31  
32 together with a large Criegee radical (reaction a3). This biradical can either stabilize (by gas  
33  
34 phase collisions) or decompose following the ester channel only.<sup>16</sup> Indeed, in the case of 2-  
35  
36 methylstyrene ozonolysis, the hydroperoxide channel, which consists on the isomerisation of  
37  
38 the excited Criegee *via* an H-atom shift to form an unsaturated hydroperoxide, is not  
39  
40 possible, since no hydrogen is available for rearrangement of the Criegee intermediate. This  
41  
42 pathway starts with a ring closure to form the dioxirane intermediate (reaction a4)  
43  
44 (Winterhalter et al., 2003),<sup>39</sup> and further O-O bond dissociation leads to the formation of  
45  
46 another biradical, which may evolve here to yield 2-methylphenylformate M2 (reaction a5)  
47  
48 and o-toluic acid M3 (reaction a6). Besides reaction a6, the formation of o-toluic acid can  
49  
50 also be explained by a pathway involving the oxidation of o-tolualdehyde either in the gas  
51  
52 phase or in the particulate phase. The presence of 2-(hydroxymethyl)benzoic acid (M4) and  
53  
54 2-formylbenzoic acid (M5) cannot be explained directly by ozonolysis reaction pathways and  
55  
56 requires an OH-attack on o-toluic acid. Reaction b could take place either in the gaseous or  
57  
58 the particulate phase. The observed decrease of o-toluic acid concentration in the particulate  
59  
60



1  
2 phase could result from its heterogeneous reaction with OH radicals leading to M4 and M5  
3  
4 formation but it cannot completely be ruled out that this decrease can also result from a  
5  
6 particle-to-gas transfer caused by o-toluic acid depletion in the gas-phase. In experiments  
7  
8 MeS\_I3, 1,3-dioxolane was introduced in the chamber as OH tracer. The tracer  
9  
10 concentration decrease was not significantly different from the decrease due to dilution  
11  
12 ( $5 \times 10^{-5} \text{ s}^{-1}$ ), therefore OH concentration was considered to be lower than  $10^6 \text{ molecules cm}^{-3}$ .  
13  
14 The formation of 2-methylphenol (M6) can result from the addition of an OH radical on the 2-  
15  
16 methylstyrene vinylic double bond (reaction c). Indeed, the decomposition of the so-formed  
17  
18 hydroxy-alkoxy radical would lead either to the formation of o-tolualdehyde (M1) *via* reaction  
19  
20 c1, or to the formation of 2-methylphenyl-peroxy radical, whose reaction with another peroxy  
21  
22 radical would produce 2-methylphenol (M6), *via* reaction c2. Finally, the equilibrium a7, which  
23  
24 could take place in the particulate phase (Figure 4), could be another possible route for 2-  
25  
26 methylphenol (M6) formation.  
27  
28

29 Further oxidation of aromatic products by OH radical may eventually lead to the formation of  
30  
31 a wide range of small unsaturated and dicarbonyl compounds.<sup>41</sup> As a matter of fact, they  
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33 were observed in SOA from OH-initiated oxidation of toluene although such volatile species  
34  
35 have high vapor pressures prohibiting their condensation to the particulate phase.<sup>42-44</sup> As an  
36  
37 example, using direct thermal desorption coupled to comprehensive gas chromatography–  
38  
39 time of flight mass spectrometry, Hamilton and co-workers estimated, from toluene photo-  
40  
41 oxidation mechanism in MCMv3.1, that ring opening compounds represented about 65 % of  
42  
43 the products.<sup>45</sup> They identified and quantified several tens of species produced by ring  
44  
45 cleavage pathways. All these species were estimated to account for about 10% of the total  
46  
47 aerosol mass concentration in the toluene photo-oxidation mechanism. In the present study,  
48  
49 although a probably larger fraction of the aerosol mass concentration was quantified, small  
50  
51 VOCs produced from the ring cleavage of aromatic derivatives were not detected for the  
52  
53 reasons exposed earlier (see section 3.3.).  
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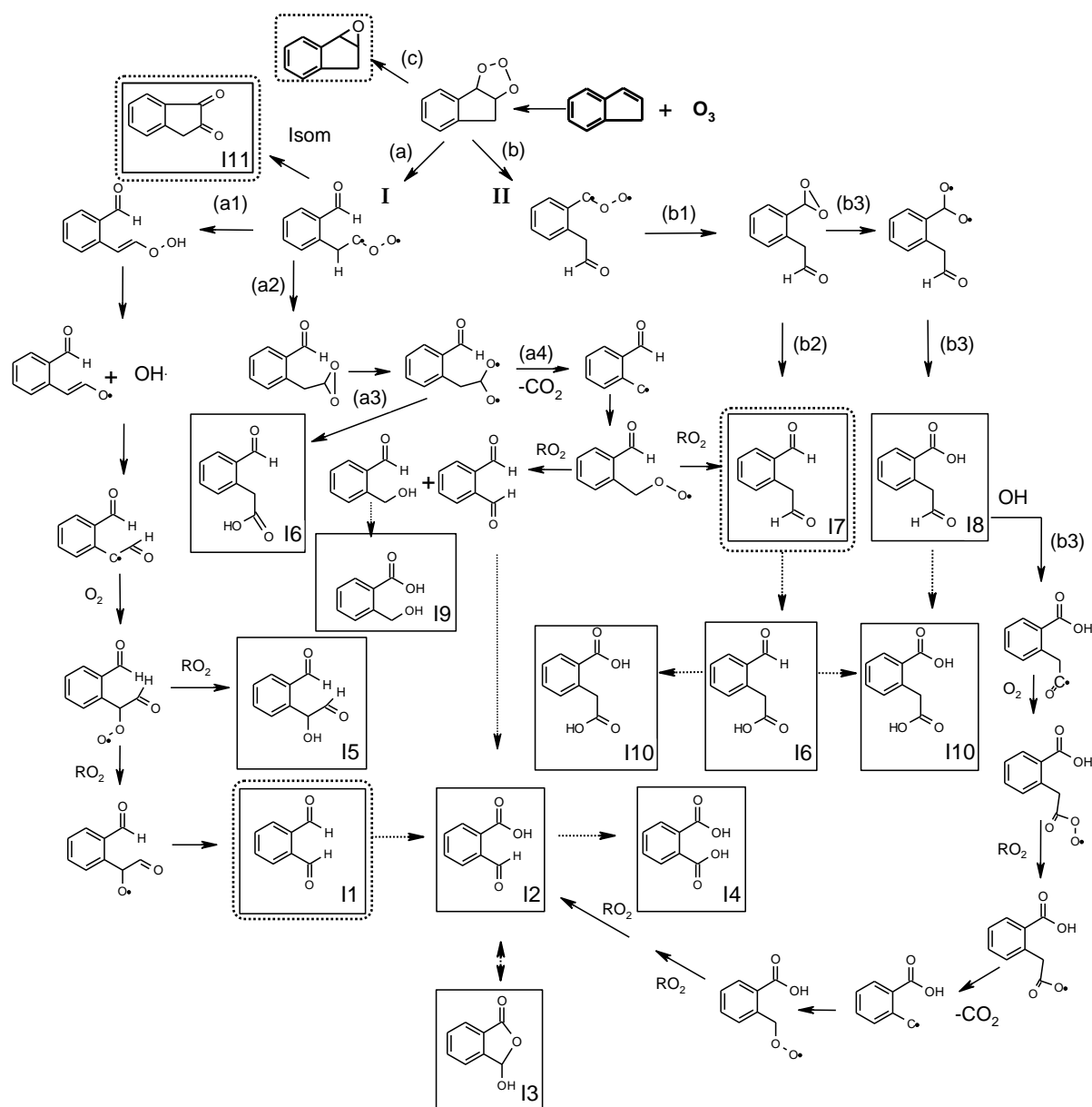


**Figure 7.** Proposed mechanism for 2-methylstyrene ozonolysis. The reactions presented with dashed lines are assumed to occur in the particulate phase. The compounds detected in the gaseous and particulate phases are respectively framed with dotted lines and with solid lines.

#### 4.2. Proposed mechanism for indene ozonolysis

Similarly to the reaction of 2-methylstyrene with ozone, reaction pathways are proposed to explain the products formed by indene ozonolysis (Figure 8). In this case, because ozone reacts on an endocyclic C=C double bond, two large Criegee radicals are formed (radicals I and II) *via* reactions a and b. These radicals can evolve by stabilization or following the ester

and the hydroperoxy channels. Indene oxide can possibly result from a minor pathway of indene ozonolysis (reaction c), as already observed previously for other alkenes.<sup>46</sup>



**Figure 8.** Proposed mechanism for indene ozonolysis. The reactions presented with dashed lines are assumed to occur within the particulate phase. The compounds detected in the gaseous and particulate phases are respectively framed with dotted lines and with solid lines.

1  
2 The Criegee radical I can only undergo the hydroperoxide channel (reaction a1), explaining  
3  
4 the formation of phthalaldehyde (I1), whose oxidation in the particulate phase can lead to  
5  
6 o-formyl benzoic acid (I2), which can isomerize to 1H-indene-1,2(3H)dione (I3) or be further  
7  
8 oxidized to form phthalic acid (I4). Moreover, before I1 formation, the peroxy radical formed  
9  
10 *via* decomposition of the hydroperoxide, can further react through self-reaction leading to 2-  
11  
12 hydroxy-2-(2-formylphenyl) acetaldehyde (I5). The Criegee radical II is not able to follow the  
13  
14 hydroperoxyde channel since no hydrogen is available to isomerize and form an unsaturated  
15  
16 hydroperoxide.  
17

18  
19 However, both Criegee radicals I and II can undergo the ester channel (reactions a2 and b1)  
20  
21 to produce two dioxiranes whose decompositions lead to intermediates. The latter can (i)  
22  
23 either directly form (2-formylphenyl) acetic acid (I6) *via* reaction a3 or lose CO<sub>2</sub> to form 2-  
24  
25 (hydroxymethyl) benzoic acid (I9) *via* reaction a4, as suggested by Winterhalter and co-  
26  
27 workers and (ii) yield 2-(formylmethyl) benzaldehyde (I7) *via* reaction b2 and 2-(2-oxo-ethyl)  
28  
29 benzoic acid (I8) *via* reaction b3.<sup>38</sup> Further oxidation of I7 and I8 leads respectively to the  
30  
31 formation of (2-formylphenyl) acetic acid (I6) and 2-carboxyphenyl acetic acid (I10).  
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33  
34 Finally, the formation of 1H-indene-1,2(3H)dione (I11) can be explained by a reaction of  
35  
36 isomerisation of Criegee radical I.  
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39 It can be noted here that phthalaldehyde (I1), produced from Criegee radical I following the  
40  
41 hydroperoxide channel, is the most abundant of the identified products in the SOA (see  
42  
43 Table 7).  
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46 Finally, the mechanism proposed here for indene ozonolysis is rather similar to the one  
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48 proposed by Riva and co-workers in their study of acenaphthylene ozonolysis.<sup>47</sup> Hence, for  
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50 each monomeric gaseous or particulate products identified in their study (except ozonide) the  
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52 corresponding product (with only one aromatic ring) was identified in the present study. This  
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54 strongly underlines the consistency between this two ozonolysis processes as logically  
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56 expected from the similarity of the alkene chemical structures.  
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#### 4.3. Comparison of SOA formation between 2-methylstyrene and indene ozonolysis

In this section, we investigate the influence of the precursor chemical structure on SOA formation. As a reminder, the major difference between the two hydrocarbon molecular structures is the nature of the reactive C=C double bond, involved in a cycle or not for indene and 2-methylstyrene, respectively. The present work is, to the best of our knowledge, the first study for mono-aromatic alkenes. Only a few studies devoted on the influence of the nature of the C=C double bond on SOA formation were reported in the literature, and these studies were focused more specifically on monoterpenes or cycloalkenes.<sup>16,17</sup>

The results presented above reveal large differences between 2-methylstyrene and indene ozonolysis in SOA formation processes and SOA chemical composition. It clearly appears from SOA mass yields (Figures 2 and 3) that indene is a much better SOA precursor than 2-methylstyrene. The yields are about 5 to 10 times higher for indene ozonolysis than for 2-methylstyrene ozonolysis, for similar consumed concentrations of precursor. Although it is not directly comparable, it is interesting to note that the same trends were observed previously. Bonn et al. compared SOA yields from  $\alpha$ -pinene (endo-cyclic double bond) and  $\beta$ -pinene (exo-cyclic double bond) and reported that  $\alpha$ -pinene ozonolysis lead to about three times higher SOA volume than  $\beta$ -pinene ozonolysis, under dry conditions and with initial hydrocarbon concentrations of about 500 ppbv.<sup>16</sup> Keywood et al. performed a whole comparative study for SOA formation yields of ten cycloalkenes.<sup>17</sup> In particular, they found that SOA mass yields obtained from the ozonolysis of methylenecyclohexane (exo-cyclic C=C double bond) were close to the ones measured for the ozonolysis of cyclohexene (endo-cyclic C=C double bond but one less carbon atom) and about twice lower than those obtained for the ozonolysis of cycloheptene (endo-cyclic C=C double bond). It should thus be highlighted that the difference in SOA mass yields between 2-methylstyrene and indene is much greater than those observed previously for non-aromatic cycloalkenes. One can assume that considering indene versus 2-methylstyrene molecular structures, there is a combination of two key features for SOA formation from ozonolysis, i.e. C=C double bond involved / not involved in a cycle and possibility / impossibility for the hydroperoxide channel, resulting in an even more favored / disfavored SOA formation. This is corroborated by recent

1 studies that reported that the hydroperoxide channel in ozonolysis opens the pathway to  
2 highly oxygenated products (HOM/ELVOC) involved in new particle formation.<sup>48</sup>  
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6 The present results should also be compared to those of Na et al. on SOA formation from  
7 styrene ozonolysis.<sup>49</sup> They measured SOA mass yields about twice higher than those  
8 obtained in our study for 2-methylstyrene. In a first approach, this result can appear  
9 inconsistent with the idea that SOA yield increase with increasing number of carbon atoms  
10 on the precursor molecular structure, for a hydrocarbon analog series. However, Keywood et  
11 al. have stressed the importance of the additional methyl group position on the molecular  
12 structure for SOA yields.<sup>17</sup> On the one hand, they measured that SOA yields were higher for  
13 1-methylcyclohexene than for cyclohexene ozonolysis, but on the other hand, they observed  
14 that 3-methylcyclohexene lead to less SOA than for cyclohexene. They proposed that this  
15 observation was due to the oligomer formation chemistry, underlining again the importance  
16 of the molecular structure on the reaction pathways and hence on SOA yields.  
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21 The distinct behaviors of 2-methylstyrene and indene can thus be explained both by their  
22 structure and by their reactivity. Indene is much more reactive toward ozone than 2-  
23 methylstyrene as shown by their reaction rate coefficients (Table 1). Therefore, under similar  
24 experimental conditions (especially the initial concentrations of reactants), semi-volatile  
25 compounds are formed and condense more rapidly in the case of indene ozonolysis.  
26  
27 Different precursor reaction rates may explain different SOA growth rates, however, it does  
28 not explain differences in the final SOA mass yields, when gas-particle equilibrium is  
29 established and SOA formation is achieved. Only the nature of the identified compounds in  
30 the particulate phase, tightly linked to the structure of the organic precursor, can help to  
31 understand the differences in the final aerosol yields. In the case of indene, ozone reacts  
32 with a C=C double bond involved in the five carbon ring, thus leading to the formation of two  
33 C9- and bi-functionalized Criegee intermediates (Figure 8), whereas in the case of 2-  
34 methylstyrene, it reacts on an external monosubstituted C=C double bond leading to the  
35 formation of only one C8- and mono-functionalized Criegee intermediate (Figure 7). It is thus  
36 likely that the first-generation products of indene ozonolysis have lower vapor pressures  
37 which favor their condensation in the particulate phase, compared to the products formed  
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2 from 2-methylstyrene ozonolysis. As an example, o-toluic acid and phthaldialdehyde were  
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4 found to be the most abundant compounds in SOA formed by 2-methylstyrene and indene  
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6 ozonolysis, respectively. Moreover, one of the two Criegee radicals formed from indene  
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8 ozonolysis undergoes the hydroperoxyde channel. This pathway seems to be particularly  
9  
10 efficient to produce low volatility species and was previously proposed to induce the  
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12 formation of dicarboxylic acids in the case of  $\beta$ -pinene ozonolysis.<sup>38,39</sup>  
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#### 17 **4. Summary and conclusions**

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19 SOA formations from 2-methylstyrene and indene ozonolysis were investigated and  
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21 compared. SOA growth curves as well as SOA final yields were measured for a wide range  
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23 of initial concentrations (from ~ 60 ppbv to ~ 1.9 ppmv). For both precursors, SOA growth  
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25 was found to be relatively linear and achieved when the hydrocarbon was consumed.  
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27 However, SOA yields were about ten times higher for indene ozonolysis than for 2-  
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29 methylstyrene ozonolysis. SOA chemical composition was studied using mass spectrometry  
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31 coupled to both on-line technique (TD-API-AMS) and off-line technique (SFE-GC-MS). For  
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33 both VOC precursors, all the identified products were ring-retaining compounds. Considering  
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35 2-methylstyrene ozonolysis, formaldehyde and o-tolualdehyde were the two main  
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37 compounds in the gaseous phase and six compounds, mainly mono-functional, were  
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39 identified in SOA : o-tolualdehyde, 2-methylphenylformate, o-toluic acid, 2-  
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41 (hydroxymethyl)benzoic acid, 2-methylbenzoic acid and 2-methylphenol. SOA chemical  
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43 composition from indene ozonolysis was found more complex and 13 products, mainly bi-  
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45 functional, were identified in the particulate phase: phthaldialdehyde, 2-formyl benzoic acid,  
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47 3-hydroxy-1(3H)-isobenzofuranone, phthalic acid, 2-hydroxy-2-(2-formylphenyl)  
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49 acetaldehyde, (2-formylphenyl) acetic acid, 2-formylmethyl)benzaldehyde, 2-(2-oxo-ethyl)  
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51 benzoic acid, 2-(hydroxymethyl) benzoic acid, 2-carboxyphenyl acetic acid and 1H-indene-  
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53 1,2(3H)dione.  
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57 Quantification of the identified components of SOA was performed and concentrations of  
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59 individual species were derived either directly from standards or from chemical surrogates. o-  
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Toluic acid and phthaldialdehyde were found to be the most abundant compounds for 2-

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2 methylstyrene and indene ozonolysis respectively. Although the identified fraction of the  
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4 aerosol mass concentration was far from being negligible (with upper limits of  $54\pm 17$  % for 2-  
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6 methylstyrene and  $25\pm 6$  % for indene), the particulate phase represented only a few % of the  
7  
8 total carbon mass balance for both precursors.  
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11 Reaction pathways were proposed to explain the identified products and discussed regarding  
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13 the precursor chemical structure. In the case of 2-methylstyrene ozonolysis, ozone reacts  
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15 with an external mono-substituted C=C double bond, which forms only one C8- and mono-  
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17 functionalized Criegee intermediate and hence more volatile products, yielding lower SOA  
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19 mass than indene under similar experimental conditions. In the case of indene, ozone reacts  
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21 with an double C=C bond involved in a five carbon ring *via* different pathways including the  
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23 hydroperoxide channel, leading to the formation of two C9- and bi-functionalized Criegee  
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25 intermediates, and hence to more highly condensable first-generation products.  
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28 Based on a large set of consistent data retrieved from experiments performed in three  
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30 different chambers and acquired using complementary analytical techniques, this work  
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32 provides relevant inputs on SOA formed by ozonolysis of aromatic alkenes and hence on the  
33  
34 contribution of anthropogenic SOA formed from the oxidation of mono- and poly-aromatic  
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36 hydrocarbon precursors. If 2-methylstyrene is obviously not a major aromatic SOA precursor,  
37  
38 it is likely that indene participates to SOA formation in polluted areas and should be taken  
39  
40 into account as an anthropogenic SOA precursor. Finally, through the detailed investigation  
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42 of the molecular composition of SOA and the comparison between the two proposed  
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44 mechanisms, this work also contributes to improve our knowledge of reaction pathways and  
45  
46 structure-reactivity relationships related to SOA formation.  
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49

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57  
58 GC-MS instrument and method development, chamber experiments and SOA molecular  
59  
60 analyses) as well as in the discussion of the results (proposition of formation pathways and



1  
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3  
4 present manuscript.  
5

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19  
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### 21 22 23 **Associated content**

24  
25 Supporting Information Available:

26  
27 Detailed procedures for gaseous phase analysis (FTIR, GC-MS, LC-MS) and particulate  
28  
29 phase analysis (TD-API-AMS, SFE-GC-MS) as well as simulation chamber experiment  
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31 procedures (ICARE, CESAM and EUPHORE facilities) and complementary figures (SOA  
32  
33 growth curves, ion signals from TD-API-AMS vs time of experiment are provided with related  
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35 references.  
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38 This material is available free of charge via the Internet at <http://pubs.acs.org>.  
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